

VIBRATIONAL SPECTRA OF BERBERINE AND THEIR INTERPRETATION BY MEANS OF DFT QUANTUM-MECHANICAL CALCULATIONS

N. BASHMAKOVA,¹ S. KUTOVYY,¹ R. ZHURAKIVSKY,² D. HOVORUN,³ V. YASHCHUK¹

¹Taras Shevchenko National University of Kyiv
(64, Volodymyrs'ka Str., Kyiv 01601, Ukraine; e-mail: *lns@univ.kiev.ua*)

²Institute of Molecular Biology and Genetics
(150, Zabolotnyi Str., Kyiv 03143, Ukraine)

³Institute of High Technologies, Taras Shevchenko National University of Kyiv
(2, bd. 5, Academician Glushkov Ave., Kyiv 03022, Ukraine)

PACS 33.20.-t, 71.15.Mb,
87.15.-v
©2011

Experimental vibrational spectra (Raman and infrared absorption) of berberine are obtained at room temperature. The vibrational spectra of berberine are calculated by the DFT method at the B3LYP/6-311++G(d,p) level. Based on the correlation between experimental and calculated data, the vibrational spectrum is interpreted in the frequency range of 800–1700 cm⁻¹ in detail. The experimental and calculated spectra of intramolecular vibrations are found to correlate closely.

1. Introduction

Natural alkaloid berberine has been used in medicine for a long time for treating a number of diseases. Preparations of berberine have been found to have cytotoxic, bactericidal, and antiviral effects [1]. Moreover, berberine influences the cancer cell metabolism by destroying cells irreversibly. Such activity of the alkaloid is attributed to its ability to intercalate into the DNA macromolecule by blocking the processes of replication and transcription. Many articles are devoted to the interaction between berberine and nuclei acids, see, e.g., reviews [1,2]. Another important target for antitumour berberine agents is DNA-topoisomerase: berberine can inhibit topoisomerase, by breaking the connection between this enzyme and DNA [3].

To clarify which mechanism causes berberine to produce its therapeutic effect, it is necessary to know the mechanisms by which it binds to DNA. Earlier, using the Raman spectroscopy method, we revealed an interesting fact of a resonance interaction of berberine and DNA vibrations [4, 5]. It was found that, in the DNA-berberine solution, the intensity of vibrations of both berberine and DNA increases greatly (by orders as compared to those of the spectra of components) in the range where

their intense spectra overlap (1000–1700 cm⁻¹). To explain this interaction, the detailed data about berberine and DNA vibrational spectra are required.

The vibrational spectra of berberine were previously investigated with the use of the Raman [6–9] and IR-absorption spectroscopy methods [7, 9], but no exhaustive interpretation of the obtained spectra was given in those works. In works [7–9], the Raman spectra of berberine in the range of 600–1800 cm⁻¹ obtained by the SERS and SSRS (surface-enhanced and shifted-subtracted Raman spectroscopy) methods were presented.

This work continues and develop our study described in [10].

2. Materials and Methods

Natural berberine is mainly contained in plants as hydrochloride or dihydrosulfate, a structural formula of a neutral berberine molecule is C₂₀H₁₉NO₅ (or C₂₀H₁₈NO₄(OH)). A structural formula of a berberine cation (as usually investigated) is [C₂₀H₁₈NO₄]⁺ (see Fig. 1). The quantum-mechanical investigations of the berberine structure were carried out in [11, 12]. The results of studies of the spatial structure of berberine cation isomers [12] demonstrate that the berberine cation that consists of hexamerous planar rings has an almost planar frame structure. A deviation from this planar structure is observed only in a partially saturated C ring. It conforms to the shape of a half-chair, in which C7 and C8 atoms of the C ring deviate appreciably from the plane of B and D rings.

The microcrystalline powder of berberine has been studied (berberine hydrochloride, “Alps Pharmaceutical”, Japan, 407 g/Mole). Since berberine absorbs in the

range to 550 nm [13], a He-Ne laser emission ($\lambda = 6328 \text{ \AA}$) has been used for the excitation of Raman spectra; the real power on a specimen was $\sim 10 \text{ mW}$. Some spectra were obtained by using a Kr-laser ($\lambda = 6471 \text{ \AA}$) with a power $\sim 60\text{--}80 \text{ mW}$ on a specimen. Raman spectra have been measured at room temperature with the 45° -reflection geometry using an optical setup made on the basis of a double-grating monochromator DFS-24 with a resolution of 1.5 cm^{-1} . A laser beam was focused on a specimen with a cylindrical lens to obtain an elongated irradiated area parallel to the entrance slit of the spectrometer. To improve the signal-to-noise ratio, the time accumulation at the point was $\sim 1.6 \text{ s}$ (multipass mode). Spectra were recorded in the range of $40\text{--}4000 \text{ cm}^{-1}$.

The IR-absorption spectra of berberine (microcrystalline powder) were recorded with a Nicolet NEXUS-470 Fourier spectrometer. We used the ATR (*Attenuated Total Reflectance*) technique. To improve the signal-to-noise ratio, a signal accumulation regime was used (128 scans). Spectra were recorded in the range of $400\text{--}4000 \text{ cm}^{-1}$ with a resolution $\leq 1 \text{ cm}^{-1}$. In the resulting spectra, Fresnel losses on the input and output surfaces of specimens were allowed for. In addition, to obtain the absorption spectra, the dependence of the penetration depth of evanescent waves into the sample on their wavelength is considered according to the relation [14]

$$d_p = \frac{\lambda}{2\pi n \sqrt{\sin^2 \theta - (n/n_s)^2}}, \quad (1)$$

where n and n_s are refractive indices of a prism (diamond) and a specimen, respectively, λ is a light wavelength, and θ is an incident beam angle.

In some cases, experimental spectra were treated, by using the *Origin* and *PeakFit* programs to specify the data on the band position, the number of band components, and so on.

The optimized geometry of the berberine cation was calculated by the DFT method at the B3LYP/6-311++G(d,p) level without any structural restrictions. At this level, the vibrational (Raman and IR-absorption) spectra were calculated in the harmonic approximation. It should be noted that, for the first time in an optimized structure, the intermolecular hydrogen bond CH...O between methoxy groups OCH_3 of berberine was found [10]. The existence of the bond was determined based on the presence of a critical point of (3,-1) type on the electron density distribution by the AIM method [15]. The binding energy was found to be 3.12 kcal/mole . The DFT calculations were performed using the “*Gaus-*

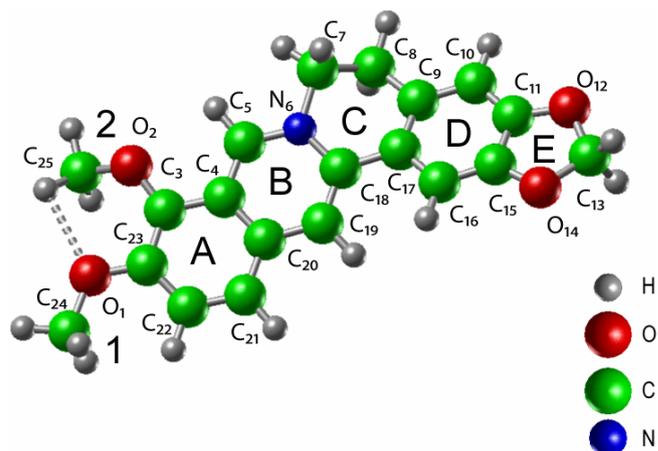


Fig. 1. Structural formula of a berberine cation (calculation by the DFT method at the B3LYP/6-311++G(d,p) level)

sian03” program package for *Win32* [16] granted by the “*Gaussian*” corporation.

Because the DFT method usually overestimates frequencies, a scaling factor of 0.985 was used for the comparison of calculated and experimental data. This correction was necessary due to errors in the calculation of interatomic interactions as a result of the limited basic functions set [17]. The value 0.985 provides the best correspondence between calculated and experimental data. The scaling factor was determined by Raman spectra – as it turned out, the shapes of calculated and experimental spectra was more similar for Raman spectra, and the correspondence between calculated and experimental frequencies was practically unambiguous. In our case, the scaling factor is close to 1 indicating that the used basic set is sufficient for the studied molecule.

In addition, it is known that the *Gaussian* set of programs deals with Raman activities S_i , not intensities I_i [18, 19]. Therefore for comparing with the experiment, the Raman activities S_i were corrected by the following relationship to obtain Raman intensities [18]:

$$I_i = \frac{C(\nu_0 - \nu_i)^4 S_i}{\nu_i [1 - \exp(-hc\nu_i/kT)]}. \quad (2)$$

Here, C is a constant, ν_0 is the laser excitation line frequency, ν_i is the vibrational frequency, h, k, c, T are Planck and Boltzmann constants, speed of light, and temperature in degrees Kelvin, respectively.

The values of frequencies and relative intensities of spectral lines with regard for the aforementioned correc-

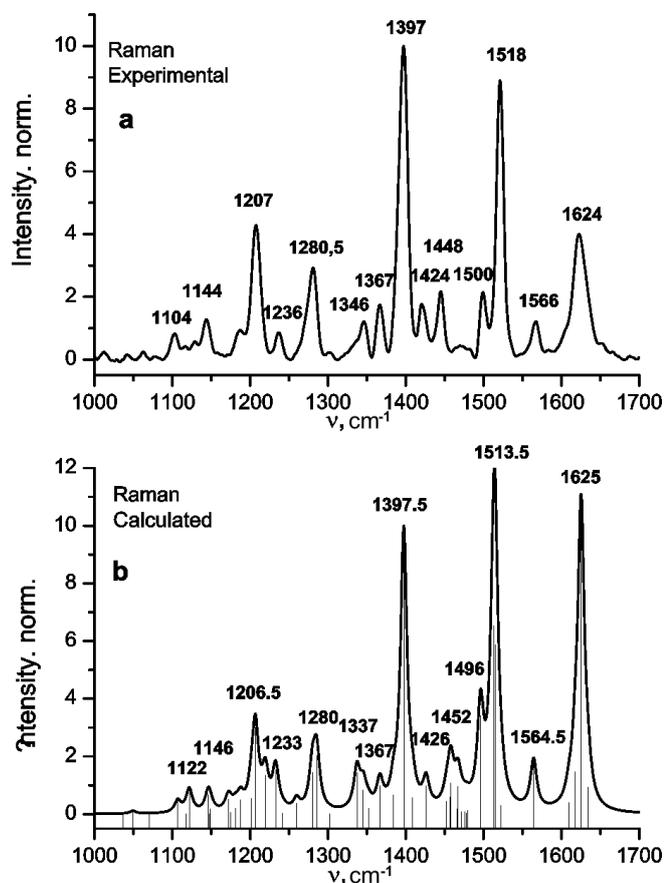


Fig. 2. Raman spectra of berberine in the range of 1000–1700 cm^{-1} : experimental, $\lambda_{\text{ex}} = 6328 \text{ \AA}$ (a), and calculated (b)

tions and the detailed interpretation of vibrations are presented in Table.

3. Results and Discussion

The berberine crystal ($\text{C}_{20}\text{H}_{18}\text{NO}_4^+\text{Cl}^- \cdot 4\text{H}_2\text{O}$) symmetry is triclinic ($P1, z = 2$) [20]. Due to a low symmetry, there are no degenerated vibrations in vibrational spectra of berberine. The correspondence between the “crystalline” and “molecular” modes is unambiguous. The most essential difference between the spectra of free and crystalline berberine had to reveal itself in the range of very low frequencies (the so-called external vibrations, there are three of them at $z = 2$). In the range of middle and high frequencies (internal vibrations), the influence of a crystal structure is very weak. Both duplication in the number of vibrations (because of $z = 2$) and some shifting of the vibrational bands take place.

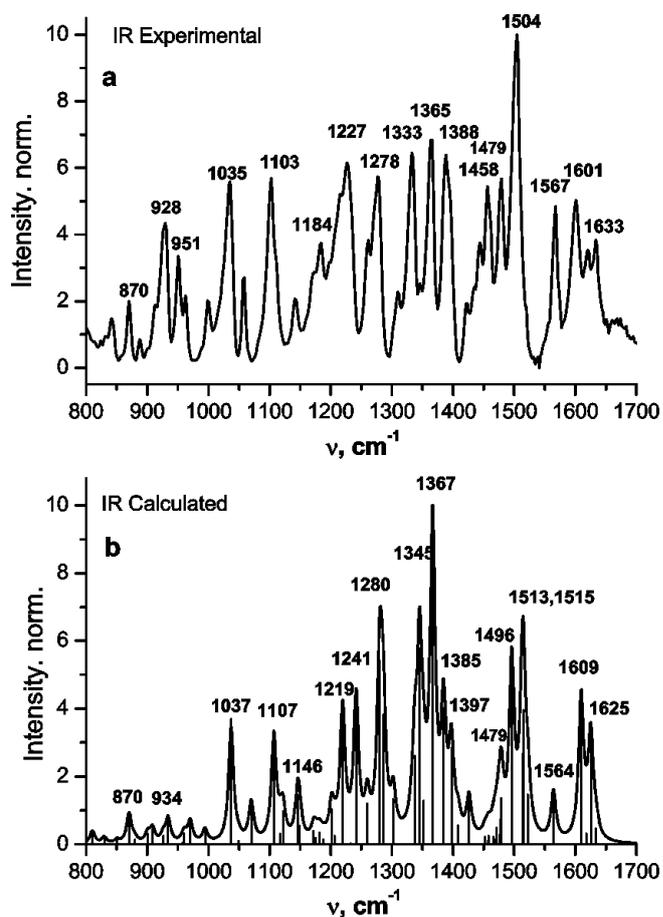


Fig. 3. IR-absorption spectrum of berberine in the range of 800–1700 cm^{-1} : experimental (a), and calculated (b)

Therefore, we believe that the comparison of experimental vibrational spectra of microcrystalline berberine with calculated spectra of the berberine cation in the actual range of 600–1800 cm^{-1} was sufficiently correct, which was confirmed by the following analysis. Experimental and calculated Raman and IR-absorption spectra in the range of 800–1700 cm^{-1} are presented in Figs. 2 and 3, respectively. The frequency values – both experimental and calculated using the scaling factor – are presented in Table. The relative intensities of Raman lines were corrected by (1). In this case, we are interested more in the range of 1000–1700 cm^{-1} . This is because berberine vibrations are the most intense in this range, and there a resonance interaction of the berberine and DNA vibrations (increasing by orders) is observed in the Raman spectrum of a berberine-DNA water solution [4, 5].

Experimental and calculated frequency values, as well as the interpretation of vibrations. The frequency scaling factor is 0.985, calculated Raman intensities were corrected by relationship (2), experimental IR-intensities – by (1)

Experimental Data					Calculations			Interpretation of vibration modes
[6] Ram	[9] Ram	[7] IR	Our data		with corrections			
			Ram	IR	ω , cm^{-1}	I_{Ram}	I_{IR}	
–	731	–	727	–	728	(2.0)	(0.1)	p(CH ₃); p, τ (CH ₂) r.C; f(CH ₂) r.E; in-plane $\nu - \delta$ of all rings; ν , δ (CNC); δ (COC);
...	–	–	–	–	–	–	–	–
–	–	–	–	vw813	810	(vw)	(0.4)	out-of-plane of atoms H,C r.A,B (H-C22,21,19); others weak
–	834	–	–	vw832	830	(0.4)	(0.2)	ν (O-CH ₃); very strong ν - δ r.D,E; weaker ν - δ r.A,B,C;
–	842	–	843	842	851	(vw)	(0.1)	out-of-plane of atoms H,C r.D (H-C10,16) –
–	–	–	–	870	870	(vw)	(1.0)	antiphase (851) and cophase (870); p(C8H2) r.C; weak ν - δ of all rings;
–	–	–	–	888	879	(0.2)	(0.1)	ν (O-CH ₃); p(CH ₂) r.C; strong ν - δ of all rings;
–	–	898	–	900	901	(vw)	(0.3)	out-of-plane of atoms H,C19 r.B (H-C19); p(CH ₂) r.C; others weak;
–	–	–	–	912	908	(0.1)	(0.5)	ν (O-CH ₃); p(CH ₂) r.C; strong O-C-O r.E; in-plane ν - δ of all rings;
–	–	–	–	928	926	(vw)	(0.3)	out-of-plane of atoms H,C5 r.B (H-C5); others very weak;
–	–	–	–	–	934	(0.2)	(0.4)	p(CH ₂) r.C; δ r.C,D; ν - δ r.E (O-C-O very strong);
–	–	–	–	–	957	(vw)	(vw)	out-of-plane of atoms H r.A,B; others very weak;
–	–	–	–	951	960	(vw)	(0.3)	δ (CH ₃ (1)); p(CH ₃ (2)); p,f(CH ₂) r.C; f(CH ₂) r.E; strong in-plane ν - δ of all rings;
–	–	–	970	963	970	(0.3)	(0.7)	weak out-of-plane r.A (C21-H, C22-H);
–	–	1002	–	1000	994	(vw)	(0.4)	p(CH ₃ (2)); ν (O1-CH ₃); f, τ (CH ₂) r.C; f(CH ₂) r.E; strong ν - δ all rings;
–	–	1035	–	1035	1037	(vw)	(3.7)	strong ν - δ (O-CH ₂ -O) r.E and ν - δ r.D; out-of-plane pf(CH ₂) r.C; weak of other rings;
–	1044	–	–	–	1050	(vw)	(0.1)	u(CH ₃); out-of-plane pf(CH ₂) r.C; f(CH ₂) r.E; ν - δ of all rings;
–	1067	1065	–	1058	1070	(vw)	(1.3)	u(CH ₃); ν (O-CH ₃); out-of-plane p(CH ₂) r.C; f(CH ₂) r.E; ν - δ of all rings;
–	–	–	1104	1103	1107	(0.4)	(3.4)	u,p(CH ₃); f(CH ₂); ν (O-CH ₃), ν (C-OCH ₂) of both pairs of bonds; in-plane ν - δ of all rings; H – in-plane;
–	–	–	–	1114sh	1117	(vw)	(0.3)	out-of-plane strong p(CH ₂) and weak of atoms O r.E; weak ν - δ of all rings;
1118	–	1110	1119	1121sh	1122	(0.8)	(1.0)	u,p(CH ₃); out-of-plane pf(CH ₂) r.E; ν (O-CH ₃), ν (C-OCH ₂), ν (N-CH ₂), ν (C7-CH ₂) r.C; δ all rings; H – in-plane and out-of-plane;
1144	–	1143	1144	1142	1146	(0.7)	(1.5)	strong ν (N-CH ₂), δ (C5NC18); ν - δ r.A,B,C and atoms H in bonds C-H; r.E rigid;
–	–	–	–	–	1148	(vw)	(vw)	strong p(CH ₃ (1)), others weak;
–	–	–	–	vw1152	1149	(0.2)	(0.6)	strong p(CH ₃ (2)), others weak;
sh	–	–	1174	1171	1172	(0.5)	(0.4)	p(CH ₃); with differ. ampl.; τ ,f(CH ₂) r.C; τ (CH ₂) r.E;
–	–	–	–	–	1175	(vw)	(0.2)	δ -vibrations r.D, mating with ν -vibrations r.C;
–	–	–	–	–	1181	(0.2)	(0.3)	δ (C7NC18); δ -vibrations r.A,B;
–	–	1182	1187.5 cm	1184	1187.5	(0.5)	(0.2)	(1181 – strong only p(CH ₃), 1187- τ (CH ₂) r.E);
–	–	–	–	1199	1201	(0.5)	(1.0)	p(CH ₃); τ (CH ₂) (r.C – strong, r.E – weak); strong ν (N-C7, N-C18) and δ (C17C9C8); weak ν of all rings;
1203	1206	–	1206	vw1207sh	1206.5	(3.0)	(0.2)	p(CH ₃); τ (CH ₂) (r.C – strong, r.E – weak); ν - δ of all rings; atoms H of rings – in-plane;

Experimental and calculated frequency values, as well as the interpretation of vibrations. The frequency scaling factor is 0.985, calculated Raman intensities were corrected by relationship (2), experimental IR-intensities – by (1) (Continuation)

Experimental Data					Calculations			Interpretation of vibration modes
[6] Ram	[9] Ram	[7] IR	Our data		with corrections			
			Ram	IR	ω, cm^{-1}	I_{Ram}	I_{IR}	
–	–	1230	1222	1227	1219.5	(1.3)	(4.2)	u(CH ₃) (1-weak); τ (CH ₂) r.C; f(CH ₂) r.E; ν (O(1.2)-C); δ (OCO) r.E; ν - δ of all rings
1235	1237	–	1236	1234sh	1233	(1.5)	(0.1)	(1219 – decreasing from A to D; 1233 – conversely);
–	–	–	–	vw1246sh	1241.5	(vw)	(4.5)	p(CH ₃); τ (CH ₂) r.C; weak f(CH ₂) r.E; ν - δ all rings and O,N;
sh	–	–	1257sh	1261	1260	(0.4)	(1.2)	weak p,u(CH ₃); τ (CH ₂) r.C; f(CH ₂) r.E; ν - δ of all rings (r.B,C,D,E – strong);
1276	1280	1271	1280.5	1278	1280	(1.4)	(5.4)	p,u(CH ₃); weak τ ,f(CH ₂) r.C; weak f(CH ₂) r.E; ν - δ of all rings (r.A,B – strong);
–	–	1301	1295	–	1285.5	(2.0)	(3.8)	p,u(CH ₃); strong f(CH ₂) r.C; weak δ (CH ₂) r.E; ν - δ of all rings (r.C,D,E – strong);
–	–	–	1303	1309	1302	(vw)	(1,3)	u(CH ₃); τ (CH ₂) r.C, weak f(CH ₂) r.E; ν - δ of all rings (r.A,B-strong);
sh	–	1331	1333	1333	1337	(1.5)	(2.6)	u(CH ₃); f,f or f, τ (CH ₂) r.C, f(CH ₂) r.E;
1342	1340	–	1346	1346	1345.5	(0.8)	(5.8)	ν - δ of all rings (1337 – strong, 1352 – weak);
–	–	–	–	–	1352	(0.2)	(1.3)	
1361	1367	1364	1367	1365	1367	(1.0)	(10)	p(CH ₃); f, τ (CH ₂) r.C, f(CH ₂) r.E;
–	–	1390	1385sh	1388	1384.5	(0.7)	(4.0)	ν - δ vibrations of all rings (1384 – commensurable, 1367 – decreasing from A to E);
1397	1397	–	1397	1397sh	1397.5	(10)	(2.8)	p(CH ₃), p(CH ₂); in-plane ν - δ -vibrations of atoms C,N,O,H of all rings;
–	–	–	1410	–	1408.5	(0.6)	(0.6)	strong f(CH ₂) r.E, weaker – r.C; weak u(CH ₃); ν - δ vibrations of all rings;
1424	1425	1424	1424	1422	1426	(1.0)	(1.3)	u(CH ₃)-strong (2); f(CH ₂); atoms O – almost immobile; strong ν - δ vibrations of all rings;
1449	1447	–	1448	1444	1452	(0.5)	(0.2)	u(CH ₃); 1457-strong (1); δ (CH ₂);
–	–	–	–	–	1457	(0.6)	(0,1)	ν - δ vibrations of all rings
–	–	–	1460	1457	1458	(1.1)	(0.2)	(1452-strong, 1457-weak), r.C,D,E (1458);
sh	–	–	1466	1466sh	1466	(0.2)	(0.2)	δ (CH ₃); 2-strong, 1-weak;
–	–	–	1472	–	1467.5	(1.0)	(0.1)	u(CH ₃); δ (CH ₂) r.C; ν - δ vibrations r.A,B,C;
–	–	–	–	–	1471	(vw)	(0.5)	
sh	–	–	1475b	–	1476	(0.1)	(0.3)	with differ. ampl. δ (CH ₃);
–	–	–	–	–	1478	(vw)	(0.7)	weak vibrations adjacent r.A;
–	1481	–	1487	1479	1479	(0.1)	(1.4)	
1501	1499	1506	1500	1504	1496	(3.4)	(5.6)	u(CH ₃); δ (CH ₂) r.E and C; ν - δ vibrations r.D (strong), r.A,B (weak);
1518	1520	–	1518	1511sh	1513.5	(6.6)	(2.7)	u, δ (CH ₃); δ (CH ₂); in-plane ν – δ -vibrations with
–	–	–	–	–	1515	(5.9)	(4.0)	differ. ampl. of atoms C,N,O,H of all rings;
–	–	–	1532sh	1520sh	1522.5	(0.3)	(1.5)	strong δ (CH ₂) r.E, weak ν -vibration r.E,D, weak of others rings;
1568	1569	1558	1566	1567	1564.5	(1.8)	(1.6)	u(CH ₃); strong (2), weak (1); strong in-plane ν - δ vibrations r.A,B (including C18-N and gr.C7H2 r.C);

Experimental and calculated frequency values as well as interpretation of vibrations. The frequency scaling factor is 0.985, calculated Raman intensities were corrected by relationship (2), experimental IR-intensities – by (1) (Continuation)

Experimental Data					Calculations			Interpretation of vibration modes
[6] Ram	[9] Ram	[7] IR	Our data		with corrections			
			Ram	IR	ω , cm^{-1}	I_{Ram}	I_{IR}	
–	–	1600	1613	1601	1609.5	(0.4)	(4.5)	in-plane ν - δ vibrations of atoms C,O,H,N of all rings;
sh	–	–	1623	–	1618	(1.5)	(0.3)	weak u,p(CH ₃); τ (CH ₂) r.C, f(CH ₂) r.E;
1626	1622	1629	1624	1620	1625	(10.7)	(3.3)	1609: strong of all rings; O – almost immobile; 1618: strong vibrations r. A,D; N – immobile; 1625: strong vibrations – r.D, N – vibrates;
sh	1633	–	1635	1633	1634	(0.9)	(0.5)	in-plane ν – δ of atoms C,N,O,H of all rings (strong – r.B,C,D, weak – gr. CH ₂ ; gr. OCH ₃ – immobile);

Notes: In brackets near frequency values are the relative intensity of lines. (The intensities of the modes at 1397 cm^{-1} in Raman spectra and 1367 cm^{-1} in IR-absorption spectra count as 10. Modes with relative intensity less than 0.1 are designated as “vw”, for “very weak”).

Abbreviations: *b*: band; *sh*: band shoulder; r.A,B,C,D,E: A, B, C, D and E rings, respectively; “weak” and “strong” denote vibrations with small and large displacements, respectively. Vibrations: ν – stretching, δ – deformation (for the CH₂-groups, the same as bending); τ : – torsion; ν – δ : a part of C atoms in a ring performs ν -vibrations, and the others perform δ -vibrations; for CH₂-groups – *f*: wagging; *p*: rocking (atoms vibrate in one plane); for CH₃-groups – *u*: umbrella (symmetric δ -vibrations); *p*: rocking (atoms vibrate in parallel planes); for τ , *f*, *p* – vibrations preserve the rigidity of CH₃ and CH₂ groups.

Unfortunately, the high frequency region of 3000–3300 cm^{-1} was not applicable to the comparative analysis: Raman spectra have very small intensities in this region, and IR-absorption spectra are only loosely correlated with calculations.

The calculated spectrum consists of 123(3 × 43 – 6) non-degenerate vibrations of a berberine cation [C₂₀H₁₈NO₄]⁺, 105 of which lie in the range of 20–1700 cm^{-1} (61 of which are in the actual range of 800–1700 cm^{-1}), other 18 ones fall in the range of 3000–3300 cm^{-1} . Low-range frequencies up to 720 cm^{-1} correspond to out-of-plane vibrations of rings and associated groups. In-plane vibrations of rings appear beginning with 720 cm^{-1} ; in particular, the only sufficiently intense 727- cm^{-1} Raman mode outside of the range of 1200–1700 cm^{-1} . In-plane vibrations are characteristic of the entire range from 720 to 1700 cm^{-1} . The high-frequency region of 3000–3300 cm^{-1} consists of the modes corresponding to vibrations of C–H-bonds with weak displacements of other atoms, mostly C.

In Fig. 2, *a*, the experimental Raman spectrum of berberine (microcrystalline powder) in the range of 1000–1700 cm^{-1} at the excitation by the 6328-Å line is presented. We adjusted the spectra, by subtracting the linearly decreased background and by smoothing the noise level slightly. In the experimental spectrum, more than twenty vibration modes were registered, and the strong correlation was observed between our data and data of [6, 9] (see Table), where the reasonably good berberine Raman spectra in the range of 1000–

1700 cm^{-1} were presented (in [9], FT-Raman). Unfortunately, in Tables in [6, 9], only the frequencies of the most intense bands were given; an interpretation of vibrations was absent, or it was incomplete or inadequate. It should be noted that the Raman spectra of berberine obtained in [7–9] by the SERS and SSRS methods are not applicable fully for the correct comparative analysis of the experimental and calculated data.

In Fig. 2, *b*, we present the calculated Raman spectrum of a berberine cation in the range of 1000–1700 cm^{-1} , by using a scaling factor.

The frequency values and the detailed interpretation are presented in Table. The strong correlation between experimental and calculated spectra is observed in the range of 1000–1700 cm^{-1} – both between frequencies and intensities of vibrations. In the range up to 1000 cm^{-1} , besides a 727 cm^{-1} mode, the intensities of calculated vibrations are very small (less by 2–4 orders) in agreement with experimental data. Notice that some modes very weak by calculations (e.g., 1302 cm^{-1}) are revealed in experimental spectra.

In Fig. 3, *a*, the IR-absorption spectrum of berberine (microcrystalline powder) in the range of 800–1700 cm^{-1} is presented (after subtracting the background), about 40 rather intense separated vibrational modes are observed, the good agreement between our data and data of [7–9] takes place (see Table).

It should be mentioned that the obtained experimental IR-absorption and Raman spectra appear quite similar to the IR-absorption spectra and the FT-Raman spec-

trum presented in [7, 9]; but IR-frequency values were presented in [7] only for the most intense lines, whereas they were not presented at all in [9]. Moreover, the analysis in [7, 9] was performed with SERS spectra that did not correspond well to our spectra by shape. The calculated Raman and IR-absorption spectra appear very similar to those we calculated, but differences in frequency values are quite substantial. Experimental and calculated Raman spectra are similar, but IR spectra are not.

We attribute this discrepancy to the fact that the calculations usually use an ion or molecule of berberine, but experiments are conducted with commercially produced crystallohydrate. Since the dipolarity of vibrations for IR-absorption processes is important, we have not discounted the possibility that the presence of dipole water molecules (and OH-groups) in specimens might have led to the dipole-dipole interaction with berberine molecules, and the results of the interaction might be revealed in IR-absorption spectra. The influence on Raman spectra is weaker, because the Raman scattering processes are related to the electron system, and the influence of the dipole-dipole interaction on the vibrational system is less direct.

In Fig. 3, *b*, we present the calculated IR-absorption spectrum of berberine (cation) in the range of 800–1700 cm^{-1} , by considering a scaling factor.

The relatively strong correlation between the frequencies of experimental and calculated IR-absorption spectra is observed in the range of 800–1700 cm^{-1} ; the correlation between the intensities of vibrations is somewhat worse, for the aforementioned reasons. It should be noted that, contrary to Raman spectra, some modes that were calculated to be intense were not observed at all in experimental IR-spectra, or they had only a very weak intensity (e.g., 1149, 1242, 1285, 1408, 1471 cm^{-1} , etc.). In this case, the lines of an unknown nature were revealed in Raman and IR spectra (1127 cm^{-1} (IR), 1129 cm^{-1} (Raman) etc.).

4. Conclusions

Thus, the Raman and IR-absorption spectra have been obtained for a berberine cation by DFT at the B3LYP/6-311++G(d,p) level. The optimized geometry of the berberine cation was calculated as well. The results of calculations strongly correlate with experimental data obtained for microcrystalline berberine chloride in the region of 1000–1700 cm^{-1} significant from the viewpoint of the interaction of berberine with DNA. The obtained interpretation of the Raman bands of berberine can be

used for an analysis of the berberine interaction with nuclei acids and other biomolecules. The DFT method can be used, with satisfactory reliability, to calculate vibrational spectra of other isoquinoline alkaloids of the protoberberine group, for which obtaining the Raman vibrational spectra would be complicated or even impossible for one reason or another.

The authors are grateful to S.O. Alekseev and M.E. Kornienko for the help in the recording of IR-spectra and to L.A. Zaika (Institute of Molecular Biology and Genetics of the NANU) for specimens of berberine.

1. L. Grycova, J. Dostal, and R. Marek, *Phytochemistry* **68**, 150 (2007).
2. M. Maiti and G.S. Kumar, *Medic. Res. Reviews* **27**, 649 (2007).
3. T.-K.E. Li, E. Bathory, E.J. La Voie, A.R. Srinivasan, W.K. Olson, R.R. Sauers, L.F. Liu, and D.S. Pilch, *Biochemistry* **39**, 7107 (2000).
4. S.Yu. Kutovyy, V.G. Pashchenko, and L.A. Zaika, *Visn. Kyiv Univ., Ser. Fiz. Mat.*, No. 7, 12 (2005).
5. V.M. Yashchuk, S.Yu. Kutovyy, N.V. Bashmakova, V.G. Pashchenko, O.V. Dudko, and L.A. Zaika, *Nauk. Zapys. Kyiv Mogyl. Akad., Ser. Fiz. Mat.*, **51**, 42 (2007).
6. I.M. Bell, R.J.H. Clark and P.J. Gibbs, *Spectrochim. Acta A* **53**, 2159 (1997).
7. N. D. Strekal, I.G. Motevich, J.W. Nowicky, and S.A. Maskevich, *J. Appl. Spectrosc.* **74**, 31 (2007).
8. S.E. Bell, E.S.O. Bourguignon, A. O'Grady, J. Villaumie, and A.C. Dennis, *Spectrosc. Eur.* **14/6**, 17 (2002).
9. M. Leona, J.R. Lombardi, *J. Raman Spectrosc.* **38**, 853 (2007).
10. N.V. Bashmakova, S. Yu. Kutovyy, D.M. Hovorun, R.O. Zhurakivsky, and V.M. Yashchuk, *Dopov. Nats. Akad. Nauk Ukr.*, No. 9, 781 (2009).
11. M.-J. Huang, K.S. Lee, and S.J. Hurley, *Int. J. Quantum Chem.*, **105**, 396 (2005).
12. V.I. Danilov, V.V. Dailidonis, D.M. Hovorun, N. Kuriita, Y. Murayama, T. Natsume, A.I. Potopalsky, and L.A. Zaika, *Chem. Phys. Lett.* **430**, 409 (2006).
13. N. Bashmakova, S. Kutovyy, V. Yashchuk, D. Hovorun, M. Losytskyy, and L. Zaika, *Ukr. J. Phys.* **54**, 471 (2009).
14. K.K. Chittur, *Biomaterials* **19**, 357 (1998).

15. F. Richard and R.F.W. Bader, *Atoms in Molecules. A Quantum Theory* (Oxford Univ. Press, Oxford, 1990).
16. Gaussian 03, Revision C.02, Gaussian, Inc., Wallingford CT, 2004.
17. M. D. Halls, J. Velkovski, and H.B. Schlegel, *Theor. Chem. Acc.* **105**, 413 (2001).
18. V. Krishnakumar, G. Keresztury, T. Sundius, and R. Ramasamy, *J. Mol. Struct.* **702**, 9 (2004).
19. D. Michalska and R. Wysokinski, *Chem. Phys. Lett.* **403**, 211 (2005).
20. B.M. Kariuki and W. Jones, *Acta Crystallogr., Sect. C: Cryst. Struct. Commun. C* **51**, 1234 (1995).

Received 01.11.10

КОЛИВАЛЬНИЙ СПЕКТР ОРГАНІЧНОЇ СПОЛУКИ
БЕРБЕРИНУ ТА ЙОГО ІНТЕРПРЕТАЦІЯ
КВАНТОВО-МЕХАНІЧНИМ МЕТОДОМ
ФУНКЦІОНАЛА ГУСТИНИ

*Н.В. Башмакова, С.Ю. Кутовий, Р.О. Жураківський,
Д.М. Говорун, В.М. Яцук*

Резюме

За кімнатної температури отримано коливальні спектри (раманівський та інфрачервоного поглинання) мікрокристалічного хлориду берберину та проведено їхню інтерпретацію методом функціонала густини на рівні теорії DFT B3LYP/6-311++G(d,p) в діапазоні частот 800–1700 cm^{-1} . Спостережено добру кореляцію між експериментальними та розрахованими частотами коливань.